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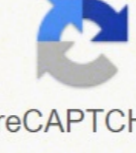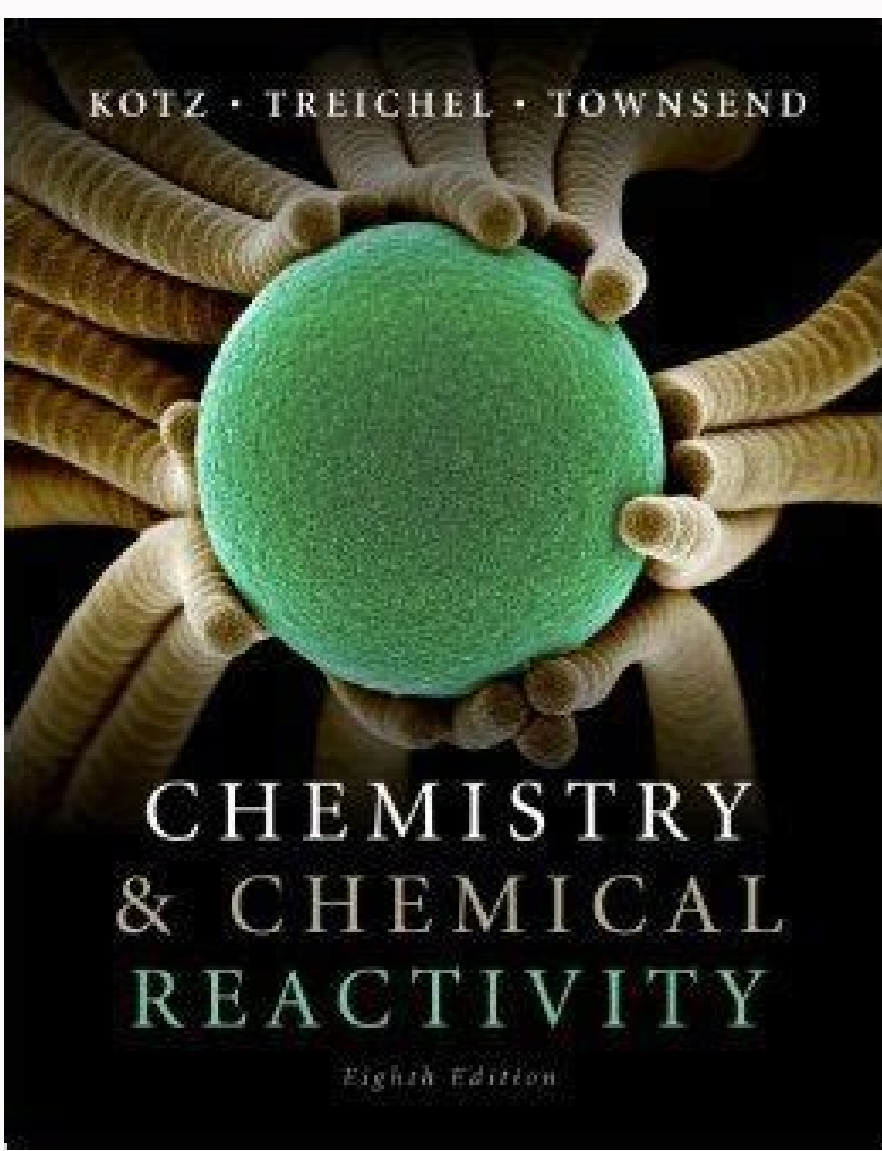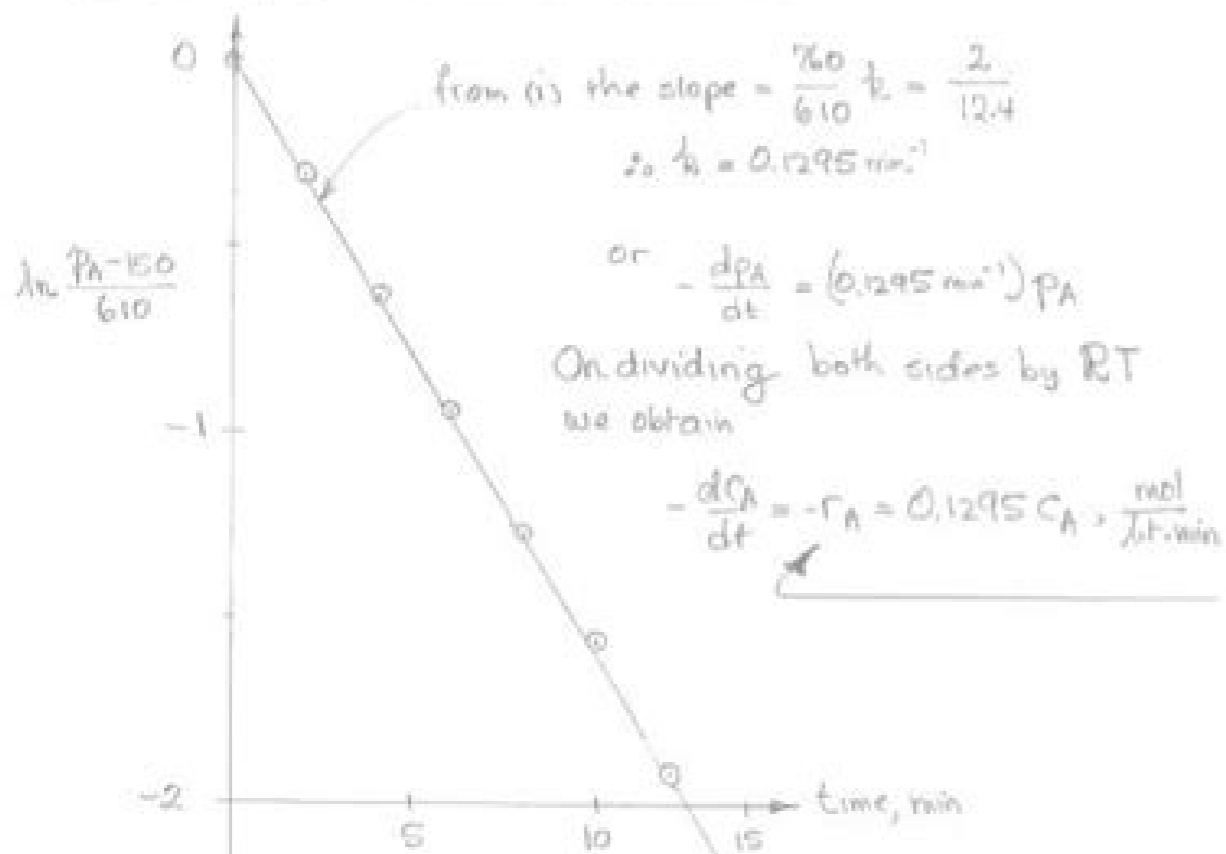
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(14)

3.25 (continued) In pressure units this integrated expression becomes

$$\ln \frac{p_A - p_{A\infty}}{p_{A0} - p_{A\infty}} = \frac{k_1 t}{(p_{A0} - p_{A\infty})/p_{A0}} \quad \cdots (i)$$
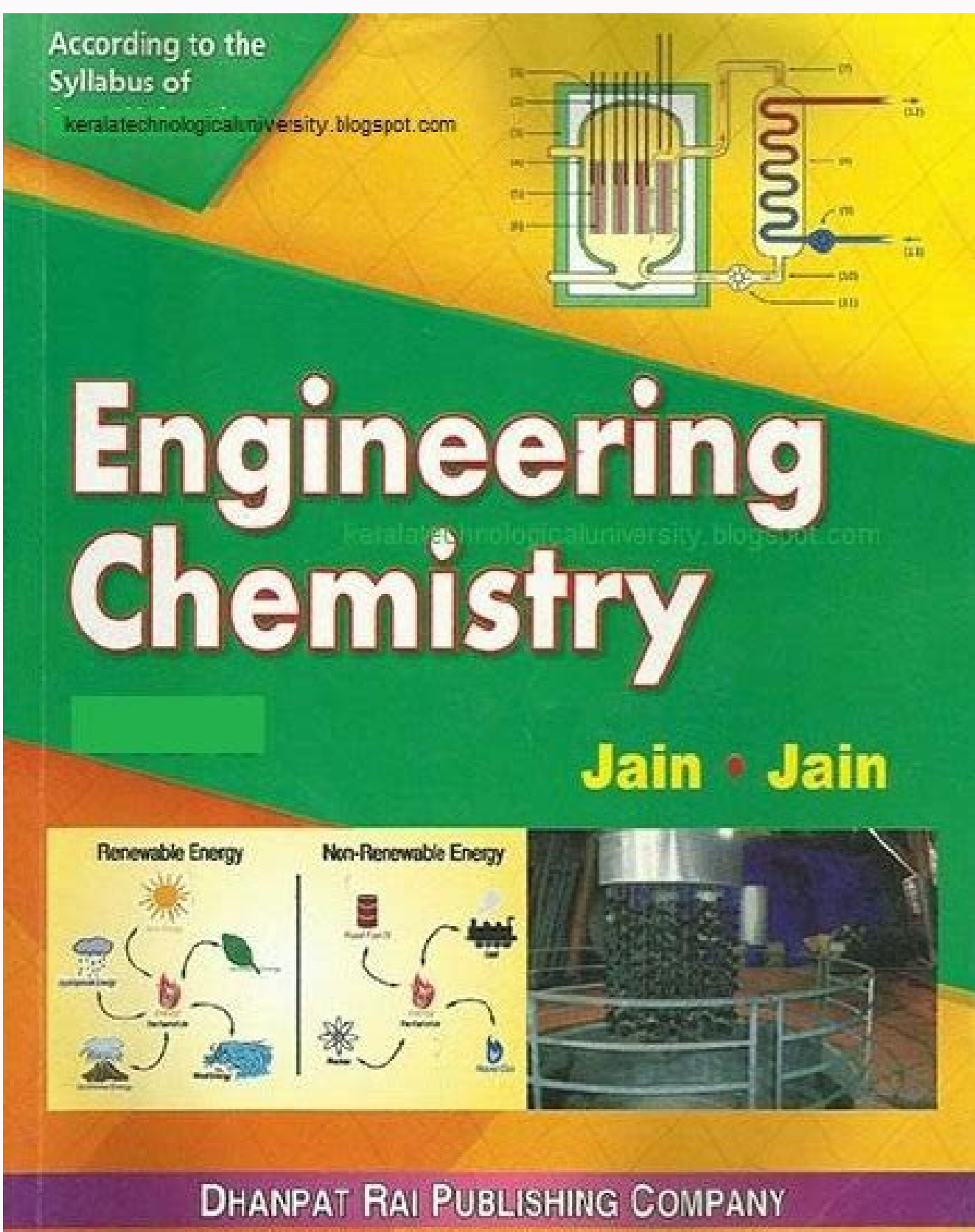
Let us see whether this equation fits the facts. So tabulate & plot

t, min	p_A	$\frac{p_A - p_{A\infty}}{p_{A0} - p_{A\infty}}$	$\ln \frac{p_A - p_{A\infty}}{p_{A0} - p_{A\infty}}$
0	760	1	0
2	600	450/610	−0.3042
4	475	325/610	−0.6286
6	390	240/610	−0.9328
8	320	170/610	−1.2777
10	275	125/610	−1.5851
12	240	90/610	−1.9136
14	215	65/610	−2.2391
∞	150	0	−∞

Note: For 1st order rxn k_c and k_p have the same value, $time^{-1}$

from (i) the slope $= \frac{760}{610} k = \frac{2}{12.4}$

so $k = 0.1295\ min^{-1}$

or $-\frac{dp_A}{dt} = (0.1295\ min^{-1})\,p_A$

On dividing both sides by RT we obtain

$$-\frac{dC_A}{dt} = -r_A = 0.1295\,C_A, \quad \frac{mol}{lit \cdot min}$$

What this shows is that if we have isothermal 1st order kinetics we can use p_A or C_A without changing the rate constants.

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